

## Naming Ionic Compounds 2

### Part 1

Write the formula for each of the following ionic compounds.

1. copper(II) acetate \_\_\_\_\_
2. sodium hydroxide \_\_\_\_\_
3. lithium oxide \_\_\_\_\_
4. cobalt(III) carbonate \_\_\_\_\_
5. aluminum sulfide \_\_\_\_\_
6. ammonium cyanide \_\_\_\_\_
7. iron(III) phosphide \_\_\_\_\_
8. vanadium(V) phosphate \_\_\_\_\_
9. sodium permanganate \_\_\_\_\_
10. manganese(III) fluoride \_\_\_\_\_
11. beryllium nitrate \_\_\_\_\_
12. nickel(III) sulfite \_\_\_\_\_
13. potassium oxide \_\_\_\_\_
14. silver bromide \_\_\_\_\_
15. zinc phosphate \_\_\_\_\_
16. copper(II) bicarbonate \_\_\_\_\_
17. nickel(II) selenide \_\_\_\_\_
18. manganese(IV) carbonate \_\_\_\_\_
19. lead(IV) nitride \_\_\_\_\_
20. tin(II) hydroxide \_\_\_\_\_

**Part 2**

Name each of the following ionic compounds. (Hint: 30 – 40 require roman numerals)

21.  $Li_3As$  \_\_\_\_\_

22.  $CaBr_2$  \_\_\_\_\_

23.  $(NH_4)_2SO_4$  \_\_\_\_\_

24.  $Al_2(CO_3)_3$  \_\_\_\_\_

25.  $AgNO_3$  \_\_\_\_\_

26.  $Mg(CH_3COO)_2$  \_\_\_\_\_

27.  $Ag_2SO_4$  \_\_\_\_\_

28.  $Zn_3N_2$  \_\_\_\_\_

29.  $NaNO_3$  \_\_\_\_\_

30.  $Cr(SO_4)_3$  \_\_\_\_\_

31.  $CuO$  \_\_\_\_\_

32.  $Pt_3(PO_4)_4$  \_\_\_\_\_

33.  $Ni(CN)_3$  \_\_\_\_\_

34.  $V_3(PO_4)_4$  \_\_\_\_\_

35.  $Co_2S_3$  \_\_\_\_\_

36.  $FeSO_3$  \_\_\_\_\_

37.  $Cu(NO_2)_2$  \_\_\_\_\_

38.  $Ni(OH)_2$  \_\_\_\_\_

39.  $Mn(NO_3)_7$  \_\_\_\_\_

40.  $Ga_2(SO_4)_3$  \_\_\_\_\_